

Chemistry Presentations

Name _____ Date _____

Group Members and phone numbers:

Due Date _____ Subject Assigned:

Objectives (Goals):

1.

2.

Grading of Presentation:

/10 Covered material in book (or packet) section completely and clearly

/5 Objectives were met

/5 Demo (presenters present) the helps understanding of section

/10 Activity (where all class members participate)

/10 Participation (will be graded by peers and teacher based upon preparation for presentation)

/5 Smooth presentation

/5 Creative and interesting

/5 Used at least 2 methods of teaching (eg. Transparencies, posters, video, outside books, white board, skit, etc.)

/5 Creative and informative handout

/60 total

Instructions: Be prepared to give a well practiced and well prepared presentation. Many students have found this activity to be one of their most favorites of the year because of the fun ideas and creativity they put in to their presentations. Make sure to have all materials set out and ready before hand and that all demos and activities are well practiced. Make sure the whole presentation (including the speaking parts) has been practiced entirely through at least a few times. Students want to have a fun, smooth, and clearly worded presentation. Make sure to get anything that needs to be copied to the teacher 1 class period before your presentation. If any chemicals are needed, get them preferably 2 class periods before the presentation. Here is a list of the group subjects:

- A. Ancient Greeks, J.J. Thomson, and Lord Kelvin
 - *Pg. 70/71 Suchocki
 - *Pg. 55/56 Zumdahl
 - *Pg. 56/57 Zumdahl Teacher's Book
- B. Dalton
 - *Pg. 77-81 Suchocki
 - *Pg. 52 Zumdahl
- C. Antoine Lavoisier
 - *Pg. 71-77 Suchocki
- D. Rutherford
 - *Pg. 87/88 Suchocki
 - *Pg. 56-59 Zumdahl
- E. Section 2.6 and page 59 in Suchocki and section 3.8 in Zumdahl.
- F. Special Assignment Group.

Notes from Class Presentations (Presenters must go over all these questions plus the ones from the atomic practice worksheet)

Group A. Ancient Greeks, J.J. Thomson, and Lord Kelvin

1. What did Aristotle think all matter was made of? Draw a picture of his model, explain it clearly, and give an example.

2. Describe Democritus's atomic model and draw a picture of it.

3. What is alchemy? And does it work?

4. What did J.J. Thomson discover and how did he discover it?

5. Describe Lord Kelvin's (his name is William Thomson but we'll call him Lord Kelvin so as to not confuse him with J.J) model and draw a picture. What can it be compared to?

Group B. Dalton

1. Write down Dalton's Atomic Theory in your own words. Draw a picture of all 5 steps.

2. Describe Dalton in his personal life.

3. Give some real life examples the 5 points in his atomic theory.

Group C. Antoine Lavoisier

1. Lavoisier knew about elements and compounds from Boyle. What are the differences between the two of them? Give an example.

2. Describe the law of conservation of mass in your own words. Give some real life examples.
3. Draw a picture of the two tin experiments Antoine did and explain what he found in each.

Group D. Rutherford

1. Draw a picture of Rutherford's gold-foil experiment. Explain how the setup worked.

2. What did Rutherford expect to happen in his experiment? Why did he think this? Instead, what happened?
3. What two important findings did Rutherford discover in his experiment?

4. Draw a picture of what Rutherford thought the atom looked like (his atomic model) and describe it.

5. How does the arm wrestling tournament represent what Rutherford found out about the positive and negative charge in an atom?

Group E. The Periodic Table (section 2.6 and 3.8)

1. On your periodic table, lightly color the metals, nonmetals, and metalloids different colors. Write down the properties of:

Metals-

Metalloids-

Nonmetals-

2. Show how to tell the number of protons, neutrons, and electrons on the periodic table. Label the atomic number and mass number. Show the example using lithium.

3. Label the alkali metals, alkaline earth metals, transition metals, halogens, and noble gases on your periodic table.

Group F. Special Assignment Group

1. Chemistry application and drawings from:

Day 1-

Day 2-

Day 3-