

3. Each color of bead stands for a particular subatomic particle (electron, proton, or neutron). Use the periodic table to figure out which color bead stands for which subatomic particle.
 - a. Blue=
 - b. Clear=
 - c. Red=

4. Now examine the unknowns A-F. For each bag, determine the number of each kind of subatomic particle, and then identify the atom in the bag.

Unknown	Number of Protons	Number of Neutrons	Number of Electrons	Identity of Atom
A				
B				
C				
D				
E				
F				

Analysis:

1. In terms of size, charge, and number, what similarities are there between protons and neutrons? What differences are there?
 - a. Size
 - b. Charge
 - c. Numbers in an atom

2. In terms of size, charge, and number, what similarities are there between protons and electrons? What differences are there?
 - a. Size
 - b. Charge
 - c. Numbers in an atom
3. In terms of size, charge, and number, what similarities are there between electrons and neutrons? What differences are there?
 - a. Size
 - b. Charge
 - c. Numbers in an atom
4. Which types of particles are the largest, contributing the most mass to an atom?
5. Use the periodic table to answer the following:
 - a. Which number on the periodic table tells you the number of protons in an atom?
 - b. Which number on the periodic table tells you the number of electrons in an atom?
 - c. How do you find the number of neutrons?
6. Which subatomic particle identifies an atom as a certain element? Justify your choice.